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### **Chapter 9 - 10 Practice Quiz**

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chemical equilibrium #Class 12

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Chemistry/Chapter#09/Chemical

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Quantities in Chemical Reactions Chapter

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10 (Liquids and Solids) - Part 1 Organic Chemistry Practice Exam 9 Testing for Halide Ions, Experiment and Explanations. A-Level Chemistry Practical 10th Class Chemistry, ch 9, Exercise Long Question Answer - Matric Part 2 Chemistry Anatomy & Physiology Chapter 9 Part A Lecture : Muscles and Muscle Tissue Chemistry 10th - Chapter 9 : Chemical Equilibrium - MCQ and Smart Syllabus 2021 Board Exam Chapter 9 Molecular Geometry and Bonding Theories Coordination Compounds | 12th Board Exam | Theory + Imp. Q & A | In 1 Video | By Arvind Arora 10th Class Chemistry, ch 9, Reversible Reaction & Dynamic Equilibrium - Matric Part 2 Chemistry Matric part 1 Chemistry, Chemistry Ch no 2 Exercise - Ch 2 Structure of an Atom - 9th Class Chapter 9 (Covalent Bonding: Orbitals) Chemistry Ch 9 Test Answers

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*MDCAT Chemistry Chapter 9 online mcq test with answers for ...*

Chapter 9 Practice Test - Naming and Writing Chemical Formulas Answer Section MATCHING 1. ANS: A PTS: 1 DIF: L1 REF: p. 253 OBJ: 9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS: H PTS: 1 DIF: L1 REF: p. 254

### *Chapter 9 Practice Test - Naming and Writing Chemical Formulas*

9.2.  $\text{FeSO}_4$  solution mixed with  $(\text{NH}_4)_2\text{SO}_4$  solution in 1 : 1 molar ratio gives the test of  $\text{Fe}^{2+}$  ion but  $\text{CuSO}_4$  solution mixed with aqueous ammonia in 1 : 4 molar ratio does not give the test of  $\text{Cu}^{2+}$  ion. Explain why? Ans: When  $\text{FeSO}_4$  and  $(\text{NH}_4)_2\text{SO}_4$  solutions are mixed in 1 : 1 molar ratio, a double salt known as Mohr's salt is formed.

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## Answers

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Chemistry is the branch of science which  
deals specifically with the elements on the  
periodic table, as well as compounds, all  
of which are composed of atoms,  
molecules, and ions. In this quiz, we aim  
to boost your knowledge on a chemistry-

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oriented level on some everyday items such as salt and aluminum foil, as well as your expertise on individual elements and their chemical symbol.

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*9th class Chemistry chapter wise online mcq test with ...*

Answer Key Testname:

CH\_08\_09\_PAC\_TEST.TST SHORT

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**ANSWER:** Write the word or phrase that best completes each statement or answers the question. 21) (a)  $DH_{rxn} = D(C''O) + 2 D(H-H) - [3 D(C-H) + D(C-O) + D(O-H)]$   
 $3 D(C-H) = -DH_{rxn} + D(C''O) + 2 D(H-H) - D(C-O) - D(O-H)$   
 $D(C-H) = (128 + 1072 + 2(436) - 358 - 463)/3 = 417$   
 $D(C-H) = 417 \text{ kJ/mol}$

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Goldsby, Kenneth Publisher McGraw-Hill  
Education ISBN 978-0-07802-151-0

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Response Practice

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9. a.  $\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$  molar  
mass of  $\text{NH}_4\text{Cl}$ , 53.49 g 0.50 mol  $\text{NH}_3$  1  
mol  $\text{NH}_4\text{Cl}$  1 mol  $\text{NH}_3 = 0.50$  mol  $\text{NH}_4\text{Cl}$   
0.50 mol  $\text{NH}_4\text{Cl}$  53.49 g  $\text{NH}_4\text{Cl}$  1 mol  
 $\text{NH}_4\text{Cl} = 27$  g  $\text{NH}_4\text{Cl}$  b.  $\text{CH}_4(\text{g}) + 4\text{S}(\text{g})$   
 $\text{CS}_2(\text{l}) + 2\text{H}_2\text{S}(\text{g})$  molar masses:  $\text{CS}_2$ ,  
76.15 g;  $\text{H}_2\text{S}$ , 34.09 g 0.50 mol S 1 mol  
 $\text{CS}_2$  4 mol S = 0.125 mol  $\text{CS}_2$  (= 0.13 mol  
 $\text{CS}_2$ ) 0.125 mol CS 2 76.15 g  $\text{CS}_2$  1 mol  
 $\text{CS}_2 = 9.5$  g  $\text{CS}_2$

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### Answers

#### *Chapter 9 Chemical Quantities*

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